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(1080p)Â.Electrochemical technology has been employed with varying degrees of success to treat drinking water. In such electrochemical approaches, the energy required to produce hydroxyl ions from water through electrical energy is often less than that required to reverse the process to decompose water into protons and electrons. The anode and cathode electrodes of an electrical water decomposition cell are, in general, highly reactive oxides or sulfides of heavy metals or mixtures thereof. Typical examples of such materials are vanadium oxides, titanium oxides, zirconium oxides, molybdenum oxides and mixtures thereof. These materials are usually supported with an electrically conductive metal base. Commonly, carbon is employed to support the reactive anode materials. Commercially available anodes are typically manufactured as from about 0.1 mm to about 5 mm diameter disks which may be disposed in a single cell or set of cells or in groups in a single cell. With existing electrochemical processes, the various reactants, such as hydrogen, oxygen and water, are continually evolved in the electrochemical cell with an overall end reaction represented by the equation: EQU 2H.sub.2 O.fwdarw.H.sub.2 +1/2O.sub.2 +2H.sup.+ +2e.sup.- (1) Although certain of the constituent materials of the electrode disks are more reactive than others, the electron flow in conventional fuel cells is generally handled by an electrochemical electrode whose major purpose is to transport electrons in the course of the oxidation-reduction reaction represented in Equation (1). In conventional electrolytic water-splitting cells, the hydrogen and oxygen gas evolved during the cell operation must be continually transferred to vessels containing the gas. It is difficult to maintain a constant gas transfer pressure in existing systems due to the fact that the hydrogen 6d1f23a050

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